

- Q.1 A student detected the pH of four unknown solution A, B, C and D as follows 11, 5, 7 and 2. Predict the nature of the solution.
- Q.2 You have four solutions A, B, C and D. The pH of solution A is 6, B is 9, C is 12 and D is 7,
- (a) Identify the most acidic and most basic solutions.
- (b) Arrange the above four solutions in the increasing order of  $H^+$  ion concentration.
- (c) State the change in colour of pH paper on dipping in solution C and D.
- Q.3 Why does 0.1 M HCL solutions have lower pH than 0.1M  $CH_3COOH$  solution?
- Q.4 Calculate the pH of 0.025M  $H_2SO_4$  solution.

