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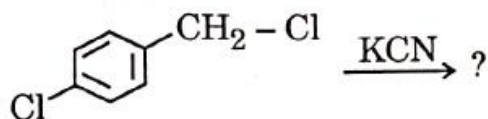## 12th CBSE CHEM SET – 1 CODE 56/3/1

1. Which of the following does *not* show variable oxidation states ?

   (A)   Fe

   (B)   Cu

   (C)   Mn

   (D)   Sc.

   (D)

2. The type of isomerism shown by the complex $[CoCl_2(en)_2]^+$ is :

   (A)   Ionisation isomerism

   (B)   Geometrical isomerism –

   (C)   Linkage isomerism

   (D)   Coordination isomerism

   (B)

3. Which of the following is diamagnetic in nature ?

   (A) Co^{3+}, octahedral complex with strong field ligand

   (B) Co^{3+}, octahedral complex with weak field ligand

   (C) Co^{3+}, in a square planar complex

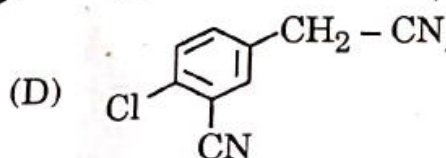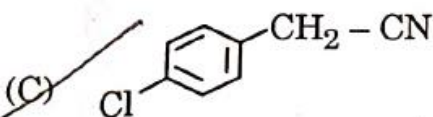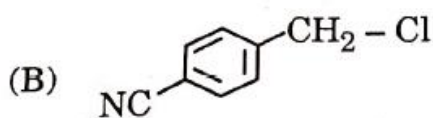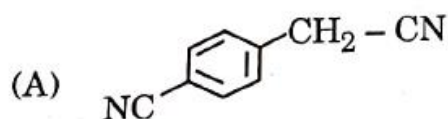   (D) Co^{3+}, in a tetrahedral complex

   [Atomic number : Co = 27]

   (A)

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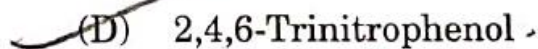**4.** Consider the following reaction :

$$\text{(4-chlorobenzyl chloride)} \xrightarrow{\text{KCN}} ?$$

The major product of the reaction is :

(A) $NC\text{—}C_6H_4\text{—}CH_2\text{—}CN$

(B) $NC\text{—}C_6H_4\text{—}CH_2\text{—}Cl$

(C) $Cl\text{—}C_6H_4\text{—}CH_2\text{—}CN$

(D) $Cl\text{—}C_6H_3(CN)\text{—}CH_2\text{—}CN$

Ⓒ

5. Which one of the following compounds has the lowest pK_a value ?

(A)   p-Cresol   (B)   p-Nitrophenol

(C)   m-Nitrophenol   (D)   2,4,6-Trinitrophenol

Ⓓ
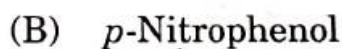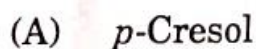
6. $(CH_3)_2CH - O - CH_3$ when treated with HI gives :

(A) $(CH_3)_2CH - I + CH_3OH$

(B) $(CH_3)_2CH - OH + CH_3 - I$

(C) $(CH_3)_2CH - I + CH_3 - I$

(D) $(CH_3)_2CH - OH + CH_3OH$

Ⓑ
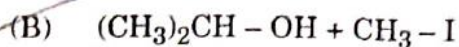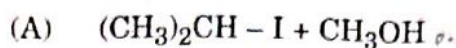
7. Which of the following compounds on treatment with benzene sulphonyl choride forms an alkali-soluble precipitate ?

   (A) CH_3CONH_2

   (B) $(CH_3)_3N$

   (C) $(CH_3)_2NH$

   (D) $CH_3CH_2NH_2$
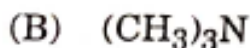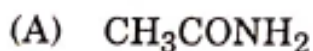
   (D)

8. The order of increasing basicities of CH_3NH_2 (I), $(CH_3)_2NH$ (II), $(CH_3)_3N$ (III) and $C_6H_5NH_2$ (IV) in aqueous media is :
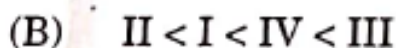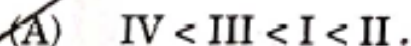
   (A) IV < III < I < II.

   (B) II < I < IV < III

   (C) I < II < III < IV

   (D) II < III < I < IV

   (A)

9. The vitamin which plays an important role in coagulating blood is :

   (A) Vitamin A

   (B) Vitamin E

   (C) Vitamin D

   (D) Vitamin K

10. When a catalyst increases the rate of a chemical reaction, then the rate constant (k) :

    (A) remains constant

    (B) decreases

    (C) increases

    (D) may increase or decrease depending on the order of the reaction

11. A 1% solution of solute 'X' is isotonic with a 6% solution of sucrose (molar mass = 342 g mol^{-1}). The molar mass of solute 'X' is :

    (A) 34·2 g mol^{-1}

    (B) 57 g mol^{-1}

    (C) 114 g mol^{-1}

    (D) 3·42 g mol^{-1}

12. During the electrolysis of aqueous NaCl, the cathodic reaction is :

    (A) Oxidation of Cl^- ion

    (B) Reduction of Na^+ ion

    (C) Oxidation of H_2O

    (D) Reduction of H_2O .

For Questions number 13 to 16, two statements are given — one labelled as Assertion (A) and the other labelled as Reason (R). Select the correct answer to these questions from the codes (A), (B), (C) and (D) as given below.

    (A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of the Assertion (A).

    (B) Both Assertion (A) and Reason (R) are true, but Reason (R) is **not** the correct explanation of the Assertion (A).

    (C) Assertion (A) is true, but Reason (R) is false.

    (D) Assertion (A) is false, but Reason (R) is true.

13. *Assertion (A) :* Addition of ethylene glycol to water lowers its freezing point.

    Reason (R) : Ethylene glycol is insoluble in water due to lack of its ability to form hydrogen bonds with water molecules.

14. *Assertion (A)*: Order of reaction and molecularity are always same for complex reactions.

   Reason (R): Order is determined experimentally and molecularity is applicable only for elementary reactions.

   (D)

15. *Assertion (A)*: The boiling point of ethanol is higher than that of dimethyl ether.

   Reason (R): Ethanol molecules are associated through hydrogen bonding whereas in dimethyl ether, it is not possible.

   (A)

16. *Assertion (A)*: Aniline does not undergo Friedel-Crafts reaction.

   Reason (R): Friedel-Crafts reaction is an electrophilic substitution reaction.

   (B)

17. (a) Define molal depression constant. How is it related to enthalpy of fusion? 2

   OR

   (b) What type of deviation is shown by ethanol and acetone mixture? Give reason. What type of azeotropic mixture is formed by that deviation? 2

Ans. (17)

(a) **Molal depression constant** -

The depression in freezing point when one mole of the solute is dissolved in 1000 gm. of the solvent is called molal depression constant.

- It is expressed as K_f.

- The relation between molal depression constant (K_f) and enthalpy of fusion (ΔH) is given by the following equation -

$$K_f = \frac{M_A \cdot RT_f^2}{\Delta H_{fus.} \times 1000}$$

Here - M_A = Molar mass of solvent

T_f = Freezing point of solvent

ΔH_f = enthalpy of fusion.

K_f = molal depression constant

R = universal gas constant

(b) Mixture of ethanol and acetone shows positive deviation from Raoult's Law.

In pure ethanol, molecules are hydrogen bonded. on adding acetone, its molecule get in between the host molecule and break some of the hydrogen bonds between them.

- Upon mixing ethanol and acetone, they form minimum boiling azeotrope mixture.

18. (a) In a reaction, if the concentration of reactant 'X' is tripled, the rate of reaction becomes twenty-seven times. What is the order of the reaction?

(b) State a condition under which a bimolecular reaction is kinetically a first-order reaction. Give an example of such a reaction. 1+1=2

2017. (18)

(a) Suppose the rate law for the given reaction is -

$$rate = K[A]^x$$

then.

$$\frac{r\, initial}{r\, final} = \frac{K[A_1]^x}{K[A_2]^x}$$

$$\frac{r}{27\,r} = \left[\frac{A}{3A}\right]^x$$

$$\frac{1}{27} = \left[\frac{1}{3}\right]^x$$

or $27 = (3)^x$

or $(3)^3 = (3)^x$

or $\boxed{x = 3}$

the order of reaction will be third.

So, the order of reaction will be third order.

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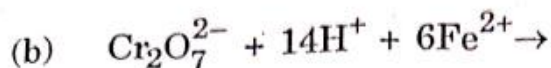JEE | NEET | 8th – 12th | Commerce | CA - Foundation | Humanities

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(b) A bimolecular reaction can be a first order reaction Kinetically, if one of the reactant is taken in access. In this condition the. concentration of the reactant, which has taken in access doesn't affect the rate of reaction.

eg. Acid catalysed hydrolysis of ethyl acetate. (Pseudo first order reaction)

$$CH_3-\overset{O}{\overset{||}{C}}-O\,C_2H_5- + H_2O_{(excess)} \xrightarrow{H^{\oplus}} CH_3-\overset{O}{\overset{||}{C}}-OH + C_2H_5OH$$
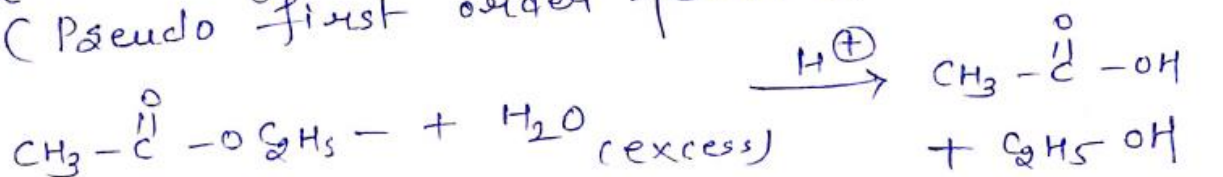
19. Complete the following ionic equations : $2\times1=2$
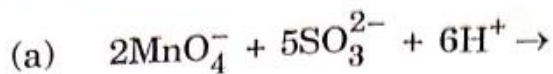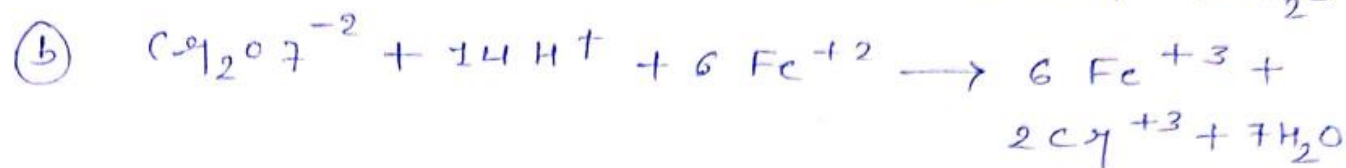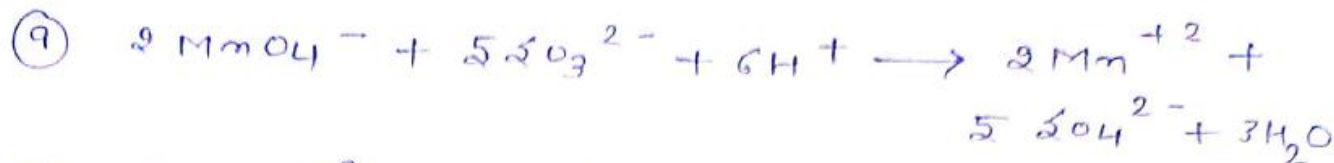
(a) $2MnO_4^- + 5SO_3^{2-} + 6H^+ \rightarrow$

(b) $Cr_2O_7^{2-} + 14H^+ + 6Fe^{2+} \rightarrow$

Sol. (19)

(a) $2MnO_4^- + 5SO_3^{2-} + 6H^+ \longrightarrow 2Mn^{+2} + 5SO_4^{2-} + 3H_2O$

(b) $Cr_2O_7^{-2} + 14H^+ + 6Fe^{+2} \longrightarrow 6Fe^{+3} + 2Cr^{+3} + 7H_2O$

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20. (a) Which halogen compound in the following pair will react faster in S_N2 reactions and why ?

$CH_3 - CH_2 - I$ OR $CH_3 - CH_2 - Br$

(b) Why is chloroform stored in closed dark coloured bottles ?   1+1=2

Sol. (20)

(a) $CH_3 - CH_2 - I$ will react faster than $CH_3 - CH_2 - Br$, because $I^\ominus$ is good leaving group than $Br^\ominus$. [Weak base are good leaving group].

(b) Chloroform oxidised in presence of air and light and form toxic phosgene gas.

21. Give reaction of glucose with the following :   2×1=2
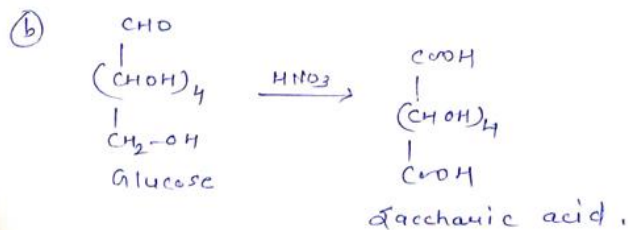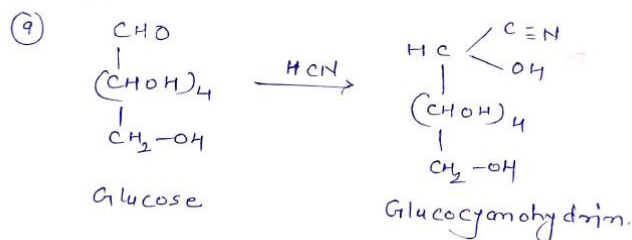
(a) HCN

(b) Conc. HNO_3

Sol. (21)

(a)
CHO
|
(CHOH)₄
|
CH₂—OH

Glucose

$\xrightarrow{HCN}$

HC
C≡N
OH
|
(CHOH)₄
|
CH₂—OH

Glucocyanohydrin.

(b)
CHO
|
(CHOH)₄
|
CH₂—OH

Glucose

$\xrightarrow{HNO_3}$

COOH
|
(CHOH)₄
|
COOH

Saccharic acid.

22. A solution is prepared by dissolving 5 g of a non-volatile solute in 200 g of water. It has a vapour pressure of 31·84 mm Hg at 300 K. Calculate the molar mass of the solute. **3**

(Vapour pressure of pure water at 300 K = 32 mm Hg)

Sol. (22) According to Raoult's law –

$$\frac{P^o - P_s}{P^o} = \frac{W_B}{M_B} \times \frac{M_A}{W_A}$$

$$\frac{32 - 31\cdot84}{32} = \frac{\cdot5}{M_B} \times \frac{18}{200}$$

$$M_B = 90$$

23. The conductivity of 0·2 M solution of KCl is $2\cdot48 \times 10^{-2}$ S cm^{-1}. Calculate its molar conductivity and degree of dissociation (α). **3**

Given :

$$\lambda^o_{K^+} = 73\cdot5 \text{ S cm}^2 \text{ mol}^{-1}$$

$$\lambda^o_{Cl^-} = 76\cdot5 \text{ S cm}^2 \text{ mol}^{-1}$$

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Sol. (23)

Molar conductivity (Λ_m) $= \dfrac{K \times 1000}{m}$

$$= \dfrac{2.48 \times 10^{-2} \times 1000}{0.2}$$

$$= 124 \quad S\ mol^{-1}cm^2$$

and $\Lambda_m^0 = \Lambda_{K^+}^0 + \Lambda_{Cl^-}^0$

$$= (73.5 + 76.5)\ S\ cm^2\ mol^{-1}$$

$$= 150\ S\ cm^2\ mol^{-1}$$

$\therefore \quad \alpha = \dfrac{\Lambda_m}{\Lambda_m^0} = \dfrac{124}{150} = 0.82$ or

$$82\%$$

24. A first-order reaction is 25% complete in 40 minutes. Calculate the value of rate constant. In what time will the reaction be 80% complete ? **3**

[Given : log 2 = 0·30, log 3 = 0·48, log 4 = 0·60, log 5 = 0·69]

Sol. (24) $K = \dfrac{2.303}{t} \log_{10}\left(\dfrac{a}{a-x}\right) \longrightarrow$ rate constant
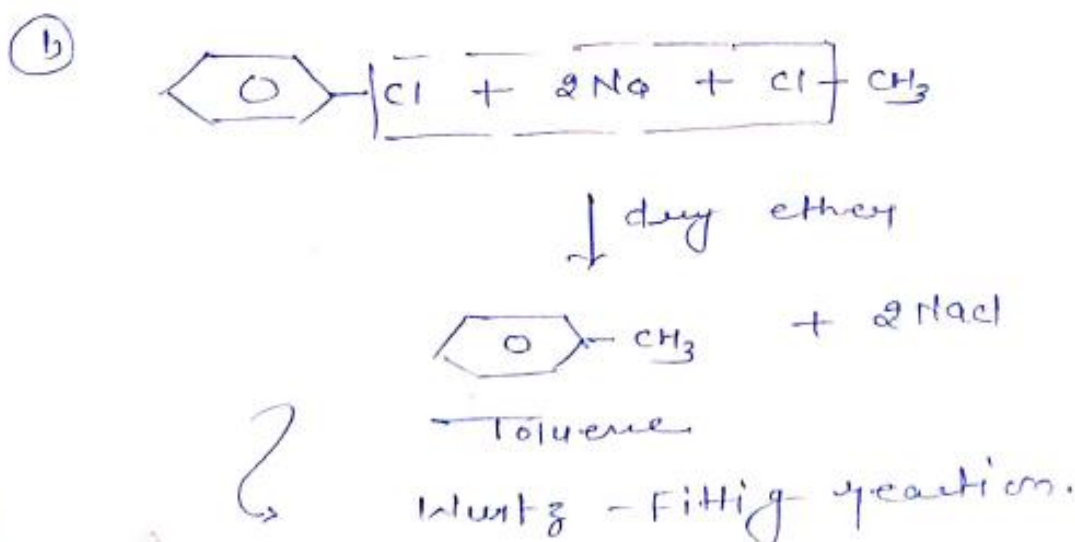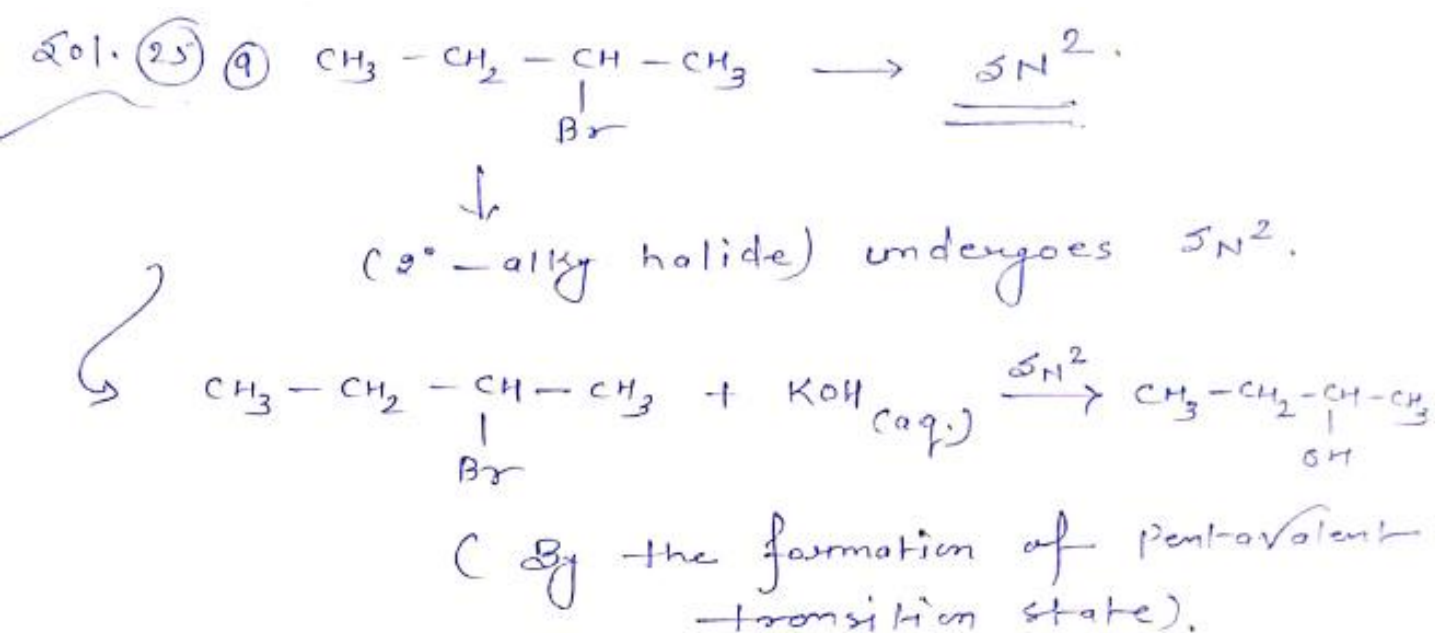
$\therefore$ K will be constant for the reaction

at for the given temp., so.

40 Min. $\leftarrow \dfrac{-t_1}{t_2} = \dfrac{\log_{10}\left(\dfrac{a}{a-x}\right)_{25\%}}{\log_{10}\left(\dfrac{a}{a-x}\right)_{80\%}}$

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25. **(a)** What type of nucleophilic substitution (S_N1 or S_N2) occurs in the hydrolysis of 2-Bromobutane to form (±)-Butan-2-ol ? Give reason.

(b) What happens when chlorobenzene and methyl chloride are treated with sodium metal in dry ether ? $\quad$ 2+1=3

Sol. (25) (a) $CH_3 - CH_2 - \underset{\underset{Br}{|}}{CH} - CH_3 \longrightarrow \underline{\underline{S N^2}}$

$\downarrow$

($2° -$ alkyl halide) undergoes S_N^2.

$CH_3 - CH_2 - \underset{\underset{Br}{|}}{CH} - CH_3 + KOH_{(aq.)} \xrightarrow{S_N^2} CH_3 - CH_2 - \underset{\underset{OH}{|}}{CH} - CH_3$

(By the formation of Pentavalent transition state).

(b) $\langle O \rangle - Cl + 2Na + Cl - CH_3$

$\downarrow$ dry ether

$\langle O \rangle - CH_3 \quad + 2NaCl$

Toluene

Wurtz – Fittig reaction.

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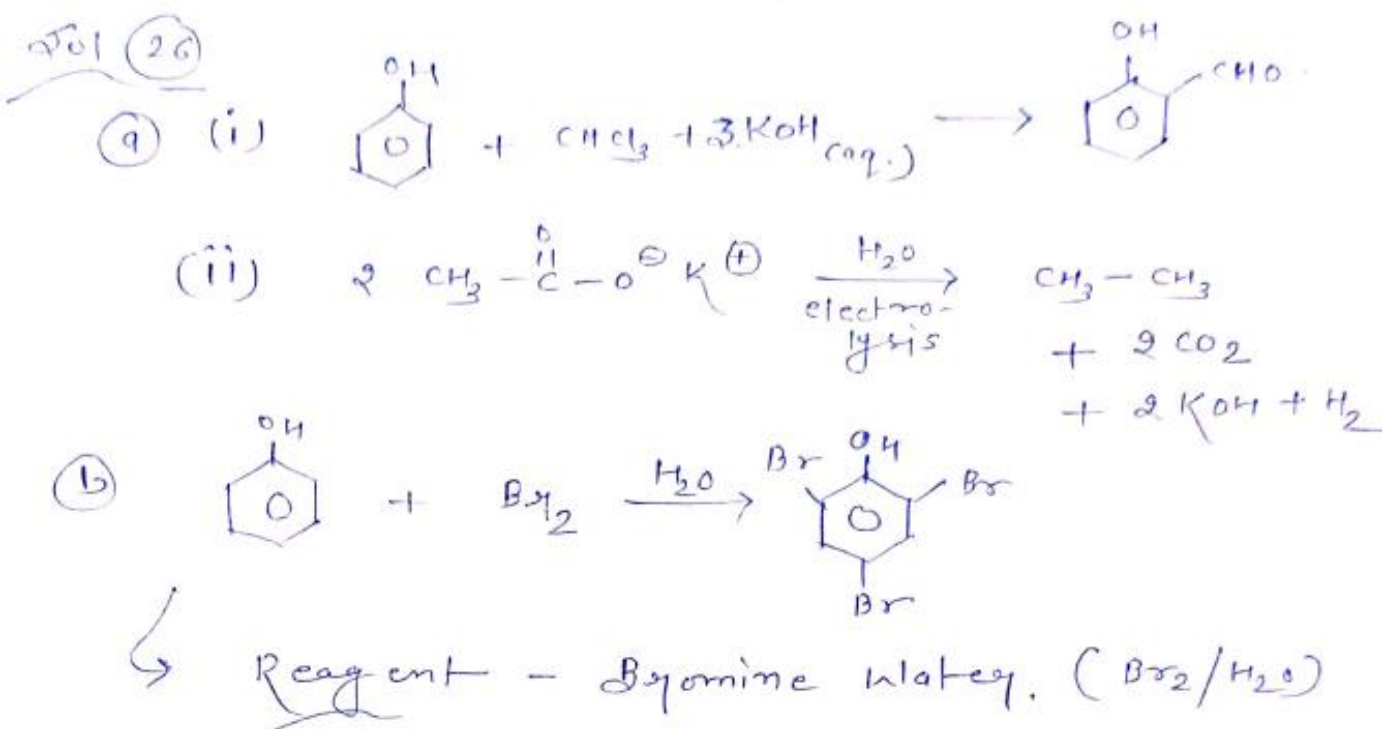JEE | NEET | 8th – 12th   Commerce | CA - Foundation   Humanities

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26. **(a)** Write the equations of the reactions involved in the following :

(i) Reimer-Tiemann reaction

(ii) Kolbe's reaction

(b) Name the reagent used in the bromination of phenol to form 2,4,6-Tribromophenol.   2+1=3

Sol (26)

(a) (i) C_6H_5OH + $CHCl_3$ + 3.KOH (aq.) $\longrightarrow$ (salicylaldehyde with OH and CHO)

(ii) $2\ CH_3-\overset{O}{\underset{}{C}}-O^{\ominus}K^{(+)}$ $\xrightarrow[\text{electrolysis}]{H_2O}$ CH_3-CH_3
$+\ 2\ CO_2$
$+\ 2\ KOH + H_2$

(b) (phenol) + Br_2 $\xrightarrow{H_2O}$ (2,4,6-tribromophenol with Br, Br, Br)

↳ Reagent – Bromine water. (Br_2/H_2O)

27. How will you bring about the following conversions ? (any **three**)   3×1=3
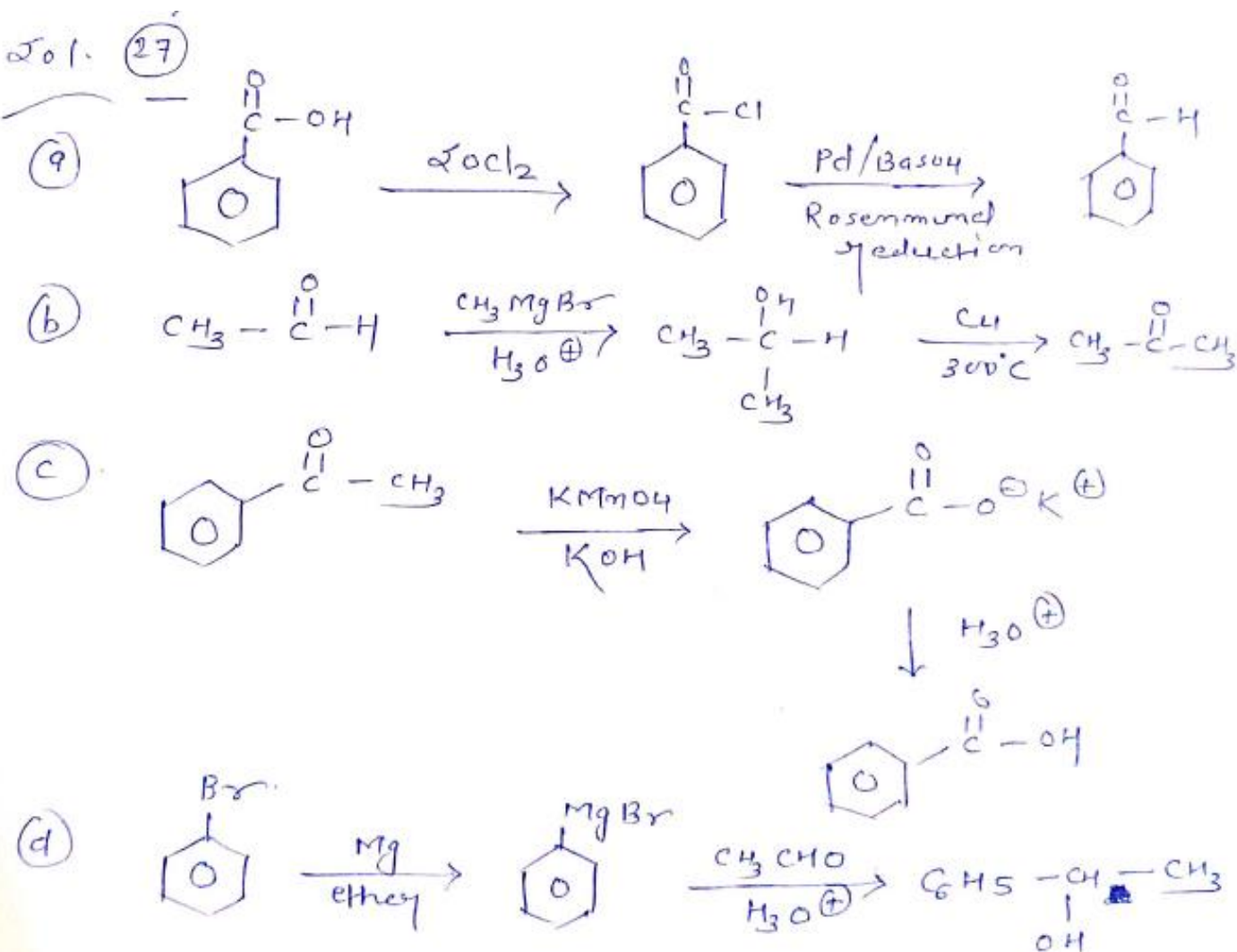
(a) Benzoic acid to Benzaldehyde

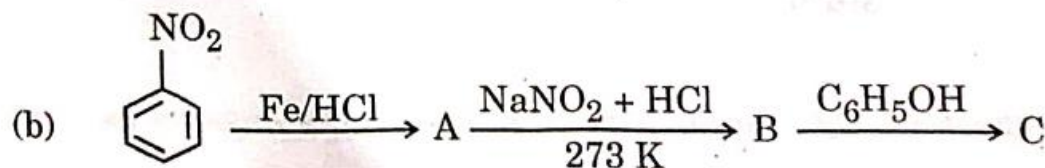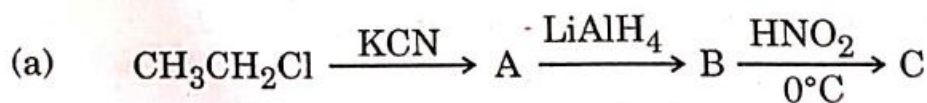**(b)** Ethanal to Propanone

(c) Acetophenone to Benzoic acid

(d) Bromobenzene to 1-Phenylethanol

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Sol. (27)

(a) Benzoic acid $\xrightarrow{SOCl_2}$ Benzoyl chloride $\xrightarrow[\text{Rosenmund reduction}]{Pd/BaSO_4}$ Benzaldehyde

(b) $CH_3-\overset{O}{\underset{}{C}}-H \xrightarrow[H_3O^{\oplus}]{CH_3MgBr} CH_3-\underset{CH_3}{\overset{OH}{\underset{|}{C}}}-H \xrightarrow[300^\circ C]{Cu} CH_3-\overset{O}{\underset{}{C}}-CH_3$

(c) $C_6H_5-\overset{O}{\underset{}{C}}-CH_3 \xrightarrow[KOH]{KMnO_4} C_6H_5-\overset{O}{\underset{}{C}}-O^{\ominus}K^{\oplus} \xrightarrow{H_3O^{\oplus}} C_6H_5-\overset{O}{\underset{}{C}}-OH$

(d) PhBr $\xrightarrow[ether]{Mg}$ PhMgBr $\xrightarrow[H_3O^{\oplus}]{CH_3CHO} C_6H_5-\underset{OH}{\overset{}{\underset{|}{C}H}}-CH_3$

28. Give the structures of A, B and C in the following reactions : $2\times1\frac{1}{2}=3$

(a) $CH_3CH_2Cl \xrightarrow{KCN} A \xrightarrow{LiAlH_4} B \xrightarrow[0^\circ C]{HNO_2} C$

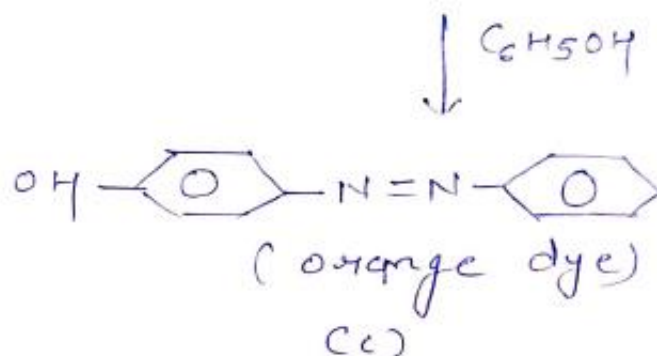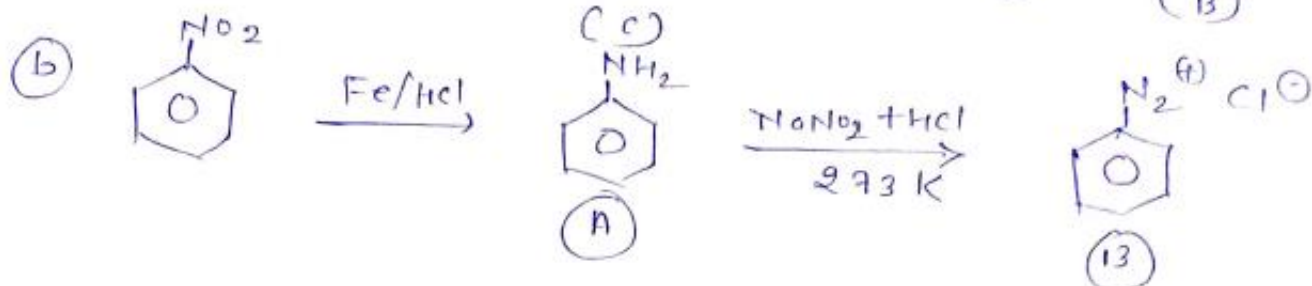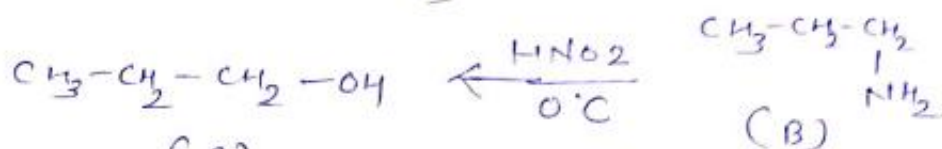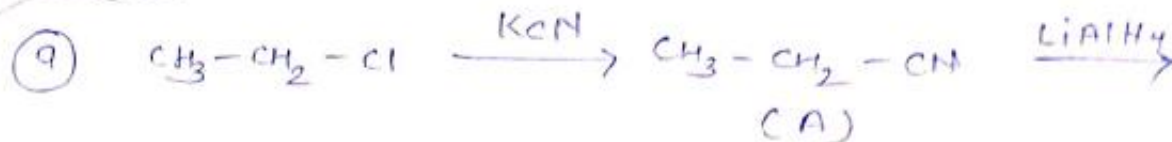(b) $C_6H_5NO_2 \xrightarrow{Fe/HCl} A \xrightarrow[273 K]{NaNO_2 + HCl} B \xrightarrow{C_6H_5OH} C$

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Solⁿ. (28)

(a) $CH_3 - CH_2 - Cl \xrightarrow{KCN} CH_3 - CH_2 - CN \xrightarrow{LiAlH_4}$
(A)

$CH_3 - CH_2 - CH_2 - OH \xleftarrow[0°C]{HNO_2}$ $CH_3 - CH_2 - CH_2 - NH_2$
(C) (B)

(b)

$NO_2 \text{(ring)} \xrightarrow{Fe/HCl}$ $NH_2 \text{(ring)} \xrightarrow[273 K]{NaNO_2 + HCl}$ $N_2^{⊕} Cl^{⊖} \text{(ring)}$
A (13)

$\downarrow C_6H_5OH$

$OH - \text{(ring)} - N = N - \text{(ring)}$
(orange dye)
(C)

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29. The involvement of $(n - 1)d$ electrons in the behaviour of transition elements impart certain distinct characteristics to these elements. Thus, in addition to variable oxidation states, they exhibit paramagnetic behaviour, catalytic properties and tendency for the formation of coloured ions. The transition metals react with a number of non-metals like oxygen, nitrogen and halogens. $KMnO_4$ and $K_2Cr_2O_7$ are common examples.

The two series of inner transition elements, lanthanoids and actinoids, constitute the f-block of the periodic table. In the lanthanoids, there is regular decrease in atomic size with increase in atomic number due to the imperfect shielding effect of 4f-orbital electrons which causes contraction.

Answer the following questions :

(a) Why do transition metals and their compounds act as good catalysts ? 　　　　　　　　　　　　　　　　　　　　　　　*1*

(b) What is the cause of contraction in the atomic size of lanthanoids ? 　*1*

(c) Define lanthanoid contraction. How does it affect the atomic radii of the third transition series and the second transition series ? 　　*2*

OR

(c) In aqueous media, which is a stronger reducing agent — Cr^{2+} or Fe^{2+} and why ? 　　　　　　　　　　　　　　　　　　　*2*

Solⁿ (29)

(a) Due to the tendency to adopt multiple oxidation states and to form complexes.

(b) Poor shielding effect of $4f$ electrons

(c) Do it yourself.

or

(c) Cr^{+2} is stronger reducing agent, because in formation of Cr^{+2} to Cr^{+3} changes is from d^4 to d^3. In d^3 electronic configuration t_{2g} orbital is half filled.

30. Proteins are the most abundant biomolecules of the living system. Proteins are the polymers of about twenty different α-amino acids which are linked by peptide bonds. Ten amino acids are called essential amino acids. In zwitter ionic form, amino acids show amphoteric behaviour as they react both with acids and bases.

On the basis of their molecular shape, proteins are classified into two types : Fibrous and Globular proteins. Structure and shape of proteins can be studied at four different levels i.e., primary, secondary, tertiary and quaternary, each level being more complex than the previous one. The secondary or tertiary structure of proteins get disturbed on change of pH or temperature and they are not able to perform their functions. This is called denaturation of proteins.

Answer the following questions :

(a)   What are essential amino acids ?   1

(b)   What is meant by zwitter ionic form of amino acids ?   1

(c)   (i)   Give one example each for Fibrous protein and Globular protein.

     (ii)   What type of linkages hold monomers of proteins together ?.   2×1=2

OR

(c)   (i)   What is the structural feature which characterises a reducing sugar ?

     (ii)   What is the structural difference between nucleoside and nucleotide ?   2×1=2

ठा01ग. (30)

(a) Do it yourself

(b) Do it yourself.

(c) (i) Fibrous protein → Kerjatin
     Globular protein → Haemoglobin.

(ii) Peptide bond.

$$OR$$

(c) (i) Do it yourself
   (ii) Do it yourself.

31. (a) (i) Calculate emf of the following cell at 25°C :

Zn (s) | Zn^{2+} (0·001M) || Cd^{2+} (0·1 M) | Cd (s)

Given : $E^o_{Zn^{2+}/Zn}$ = – 0·76 V, $E^o_{Cd^{2+}/Cd}$ = – 0·40 V [log 10 = 1]

(ii) State Faraday's second law of electrolysis. How will the pH of aqueous NaCl solution be affected when it is electrolysed ? 3+2=5

OR

(b) (i) Calculate the $\Delta_r G^o$ and log K_c for the following cell reaction :

Fe (s) + Ag^+ (aq) $\rightleftharpoons$ Fe^{2+}(aq) + Ag (s)
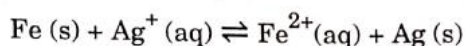
Given : $E^o_{Fe^{2+}/Fe}$ = – 0·44 V, $E^o_{Ag^+/Ag}$ = + 0·80 V,

1 F = 96500 C mol^{-1}

(ii) Write any two advantages of the fuel cells over primary and secondary batteries ?

(iii) How many Faradays are required for the oxidation of 1 mole of H_2O to O_2 ?
3+1+1=5

Sol. (31)

(a) (i)

$$E^o_{cell} = E^o_{cathode} - E^o_{anode}$$

$$= -0.40 - (-0.76)$$

$$= -0.40 + 0.76$$

$$= 0.36 \checkmark$$

$$F_{cell} = E^o_{cell} - \frac{0.059}{n} \log_{10} \frac{[P]}{[R]}$$

$$= 0.36 - \frac{0.059}{2} \log_{10} \frac{[10^{-3}]}{[10^{-1}]}$$

$$= + 0.419 \checkmark$$

(ii) Theory based.

$$\underline{OR}$$

(b) (i)

$$Fe_{(s)} \longrightarrow Fe^{+2}_{(aq.)} + 2e^-$$

$$Ag^+_{(aq.)} + e^- \longrightarrow Ag^\circ_{(s)} \quad \times 2$$

$$Fe_{(s)} + 2Ag^+_{(aq.)} \longrightarrow Fe^{+2}_{(aq.)} + 2Ag^\circ_{(s)}$$

$$\therefore \Delta_r G^\circ = -nFE^\circ_{cell}$$

$$\begin{cases} E^\circ_{cell} = E^\circ_{cathode} - \\ \qquad E^\circ_{anode} \\ = 0.80 - (-0.4 \\ = +1.24\,V. \end{cases}$$
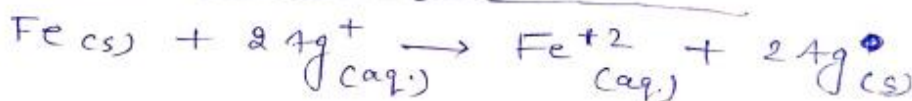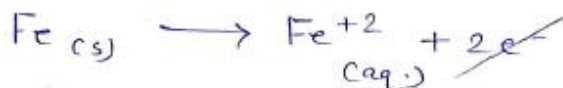
$$= -2 \times 96500 \times 1.24$$

$$= 239320 \; J/mol$$

$$= 239.32 \; KJ/mol.$$

$$\therefore \Delta G^\circ = -nFE^\circ = 239.32 \; KJ/mol.$$

$$and \quad \Delta G^\circ = -2.303 \, mRT \, log_{10} K = 239320.$$

$$\therefore K = 10^{-20.97} = 8 \times 10^{2} \not{\!\!/} \; 2 \times 10^{-21}$$

2.

(i) $E_{cell} = E^\circ - \dfrac{0.0591}{2} \, log. \dfrac{[Zn^{2+}]}{[Cd^{2+}]}$

(ii) do it.
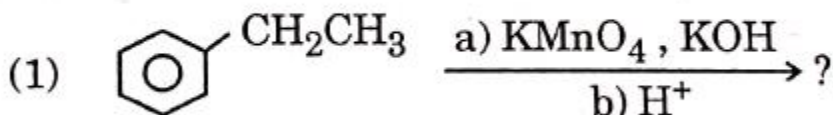
(b) (i) $\Delta G^\circ = -2.303 \, RT \, log \, K_{eq.}$

(ii) It produce pure water and Not discharge.

(iii) $2H_2O \longrightarrow O_2 + 4H^+ + 4e^-$

$\quad$ 2 mol $\qquad\qquad\qquad$ 4e⁻

$\quad$ 1 mol $\qquad\qquad\qquad$ 2e⁻ = 2F
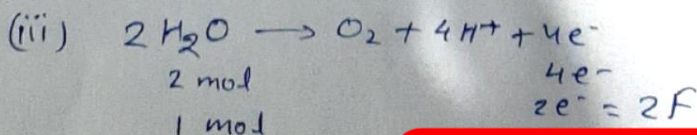
32. (a) (i) Write the major product(s) in the following reactions :

(1) CH_2CH_3 $\xrightarrow[\text{b) } H^+]{\text{a) } KMnO_4, KOH}$?

(2) $- CHO + CH_3 - \overset{\overset{\displaystyle O}{\|}}{C} - CH_3 \xrightarrow{\text{dil NaOH}}$?

(3) $\xrightarrow{Br_2 / FeBr_3}$?
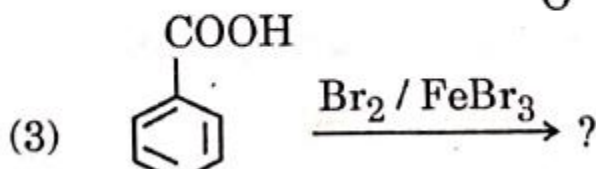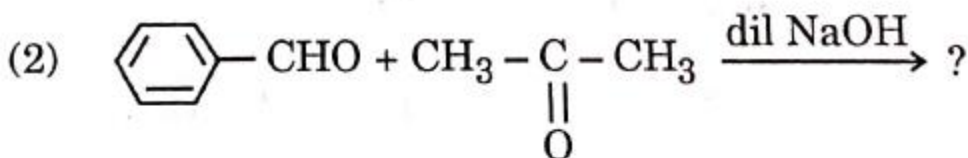
(ii) Give simple chemical tests to distinguish between the following pairs of compounds :

(1) $COCH_3$ and $COCH_2CH_3$

(2) Pentanal and Pentan-3-one $\hspace{2cm}$ 3+2=5

OR

(b) (i) Give reasons for the following :

(1) In semicarbazide, only one $- NH_2$ group is involved in the formation of semicarbazone.

(2) Acetaldehyde is more reactive than acetone towards addition of HCN.

(ii) (1) Arrange the following in decreasing order of their acidic strength :

$CH_3COOH, \quad O_2N - CH_2 - COOH, \quad HCOOH$
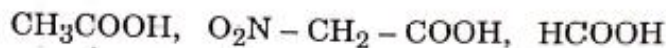
(2) Name the reagent in the following reaction :

$CH_3 - CH = CH - CH_2 - CN \xrightarrow{?}$
$\hspace{3cm} CH_3 - CH = CH - CH_2 - CHO$
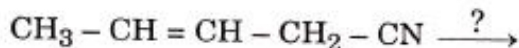
(iii) Write the reaction involved in Hell-Volhard-Zelinsky reaction.

$\hspace{8cm}$ 2+2+1=5

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12th CBSE CHEM SET – 1 CODE 56/3/1

Sol^n (32)

(a) (i)
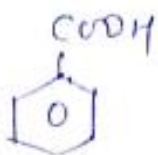
(1) benzene-COOH   (oxidation of alkyl benzene)
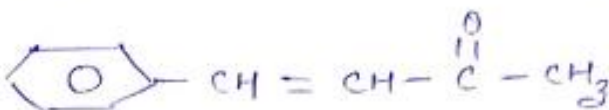
(2) benzene-$CH=CH-\overset{O}{\overset{||}{C}}-CH_3$ (Aldol condensation)
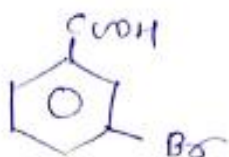
(3) benzene with COOH and Br   (—COOH is deactivating and meta directing group)

(ii)

(1) Haloform test

(2) Tollen's reagent

OR.

(b) (i)

(1) $NH_2 - NH - \overset{O}{\underset{||}{C}} - \ddot{N}H_2 \longrightarrow$ The e.p. of other NH_2 group is delocalised and involved in resonance.
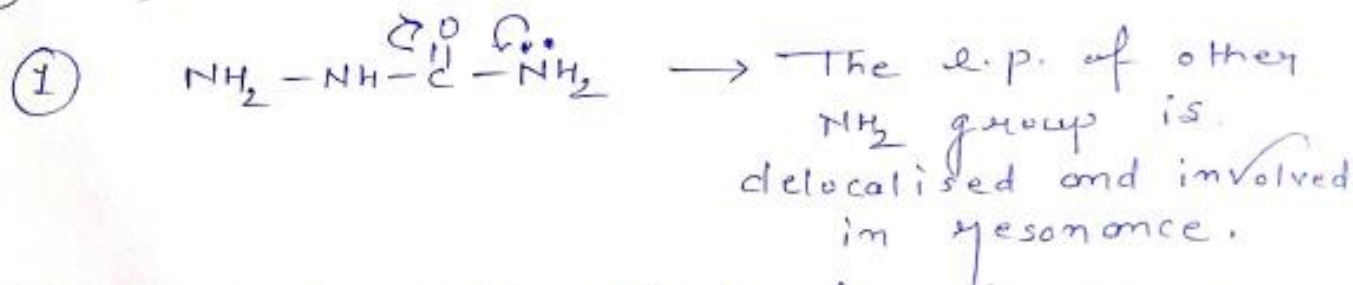
(2) Due to $+I$ effect of $-CH_3$ group.

(ii)

(1) $O_2N - CH_2 - COOH > HCOOH > CH_3 COOH.$

(2) $SnCl_2/HCl$ (Stephen's reduction).

(iii) Do it yourself.

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33. Attempt any *five* of the following : $5 \times 1 = 5$

(a) Write the IUPAC name of the complex :

$$[Co(H_2O)(CN)(en)_2]^{2+}$$

(b) Why is geometrical isomerism not possible in tetrahedral complexes having two different types of unidentate ligands coordinated with the central metal ion ?

(c) Arrange the following complex ions in increasing order of their crystal field splitting energy (Δ_0) :

$$[Co(NH_3)_6]^{3+}, \ [CoF_6]^{3-}, \ [Co(CN)_6]^{3-}$$
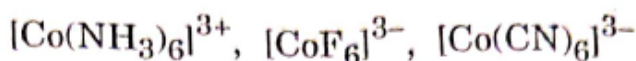
(d) Write the hybridization and magnetic character of the complex $[Ni(CO)_4]$ on the basis of valence bond theory.

[Atomic No. : Ni = 28]

(e) Out of $[CoF_6]^{3-}$ and $[Co(C_2O_4)_3]^{3-}$, which one complex is :

(i) more stable ?

(ii) the high spin complex ?

(f) What is the difference between an ambidentate ligand and bidentate ligand ?

(g) Write the electronic configuration of d^5 in terms of t_{2g} and e_g in an octahedral field when :

(i) $\Delta_0 > P$, and (ii) $\Delta_0 < P$

(a) Aqua cyanido Bis-(ethylene diammine) cobalt(II) ion

(b) Not planar.

(c) $[Co(CN)_6]^{3-} > [Co(NH_3)_6]^{3+} > [CoF_6]^{3-}$

(d) sp^3 , $\mu = \sqrt{n(n+2)}$, $n = 0$

$= 0$

(e) (i) $[Co(C_2O_4)_3]^{3-}$ $C_2O_4^{2-}$ strong ligand

(ii) $[CoF_6]^{3}$ F^- weak ligand

(f) do it.

(g) do it.

12th CBSE CHEM SET – 2 CODE 56/3/2

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2024

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1.5 Yr. Classroom	1 Yr. Classroom	2 Yr. Classroom	2 Yr. Classroom	3 Yr. Classroom

2023

98.13 (Percentile)	2169 (AIR)	96.12 (Percentile)	93.51 (Percentile)	
NIT TRICHI*	IIT KHARAGPUR	NIT RAIPUR	NIT RAIPUR	**4 / 27**
GAURAV BAIRAGI	ADITYA P. SINGH	MAHIMA SHUKLA	MAYANK K GUPTA	SELECTION RATIO
1 Yr. Classroom	4 Yr. Classroom	1 Yr. Classroom	2 Yr. Classroom	Got NIT or IIT

2022

99.44 (Percentile)	99.40 (Percentile)	98.20 (Percentile)	96.37 (Percentile)	96.12 (Percentile)
NIT SURATHKAL	NIT ROURKELA	NIT ROURKELA	NIT RAIPUR	VIT VELLORE
SANIDDHYA SINGH	SWATI KATAILIHA	KUNAL SHARMA	PRAKHAR SHRIVASTAVA	RADHIKA CHOPDE
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8 / 30 SELECTION RATIO — Got NIT or IIT

Success Stories : NEET (Last 2 Years)

MBBS 1st Attempt 2023	MBBS 1st Attempt 2023	MBBS 1st Attempt 2023	MBBS 1st Attempt 2022	MBBS 1st Attempt 2022
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4 Yr. Classroom	4 Yr. Classroom	Short Term Course	Short Term Course	3 Yr. Classroom
GMC, JAGDALPUR	GMC, KANKER	GMC, RAIPUR	GMC, KANKER	GMC, KORBA

MBBS 1st Attempt 2022	MBBS 1st Attempt 2022	MBBS 1st Attempt 2022		
ANVITAA MURTY	SIDDHI AGRAWAL	PRIYAL YADAV	SELECTION RATIO 2023	SELECTION RATIO 2022
4 Yr. Classroom	4 Yr. Classroom	4 Yr. Classroom	**03 / 43**	**05 / 24**
IMS, Bhubaneswar	YENEPOYA, MANGLORE	RIMS, RAIPUR	Got MBBS	Got MBBS